

Answers to Chapter 14

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Solve the problems below:

1. What is the molarity of a solution in which 10.0 g of NaCl are dissolved in 1.00 L of solution?

$$M = \frac{10.0 \text{ g NaCl}}{1.00 \text{ L}} = 0.171 \text{ M}$$
2. What is the molarity of a solution in which 10.0 g of AgNO₃ is dissolved in 500 mL of solution?

$$M = \frac{10.0 \text{ g AgNO}_3}{0.500 \text{ L}} = 0.171 \text{ M}$$
3. How many grams of NaCl should be used to prepare 750 mL of a 0.200 M solution?

$$g = M \times V \times \frac{m}{1000}$$

$$g = 0.200 \text{ M} \times 750 \text{ mL} \times \frac{58.44 \text{ g}}{1000 \text{ mL}} = 8.77 \text{ g}$$
4. To what volume should 1.0 g of NaCl be added in order to prepare a 0.20 M solution?

$$V = \frac{g}{M \times \frac{m}{1000}}$$

$$V = \frac{1.0 \text{ g}}{0.20 \text{ M} \times \frac{58.44 \text{ g}}{1000 \text{ mL}}} = 0.85 \text{ L}$$
5. How many grams of CaCl₂·2H₂O are needed to prepare 100 mL of a 0.10 M solution?

$$g = M \times V \times \frac{m}{1000}$$

$$g = 0.10 \text{ M} \times 100 \text{ mL} \times \frac{147.01 \text{ g}}{1000 \text{ mL}} = 1.47 \text{ g}$$

Answers to Chapter 14

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